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## Practice Sheet for Periodic Table:

Q.1 In each of the following pairs, circle the species with the higher first ionization Potential energy:

(a) Li or Cs

(b)  $\text{Cl}^-$  or Ar

Ans: \_\_\_\_\_

Ans: \_\_\_\_\_

(c) Ca or Br

(d)  $\text{Na}^+$  or Ne

Ans: \_\_\_\_\_

Ans: \_\_\_\_\_

(e) B or Be

Ans: \_\_\_\_\_

Q.2 In each of the following pairs circle the elements with the larger atomic radius:

(a) Mg or Ba

Ans: \_\_\_\_\_

b) S or S<sup>2-</sup>

Ans: \_\_\_\_\_

(c) Cu<sup>+2</sup> or Cu

Ans: \_\_\_\_\_

(d) He or H<sup>-</sup>

Ans: \_\_\_\_\_

(e) Na or Cl

Ans: \_\_\_\_\_

Q.3 Circle the best choice:

(a) Highest value of electro-negativity:

Li, Na, K

(b) Largest radius:

Mg, Ca, Sr

(c) Highest electro-negativity in the periodic table:

Li, Be, F

(d) Smallest atom:

Na, Li, Be

(e) Transition metal:

Fe, Ca, Na

(f) Lowest ionization energy:

K, Na, Ca

(g) Highest value of second ionization potential:

Na, Mg, Ba

(h) Identify the groups of the following elements:

Ar, K, Ca