

Chemistry:

Atomic Number, Mass Number & Isotopes:

1. Give the information...



Element: _____

Protons: _____

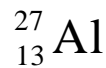
Neutrons: _____



Element: _____

Atomic #: _____

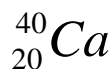
At.Mass : _____



Element: _____

Protons: _____

Electrons: _____

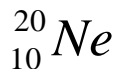


Element: _____

Protons: _____

Neutrons: _____

Atomic mass: _____



Element: _____

Atomic No: _____

Electrons: _____

Atomic mass: _____



Element: _____

Protons: _____

Electrons: _____

Neutrons: _____



Element: _____

Atomic No: _____

Mass No: _____

Electrons: _____



Element: _____

Atomic No: _____

Electrons: _____

Neutrons: _____



Element: _____

Protons: _____

Electrons: _____

Neutrons: _____

2. Write the symbol for the ion:

- A. An Oxygen Atom That Has Gained Two Electrons: _____
- B. A Magnesium Ion That Has Lost Two Electrons: _____
- C. A Copper Ion That Has Lost One Electron: _____
- D. A Phosphorous Atom That Has Gained Three Electrons: _____
- E. A Sodium Atom That Has Lost One Electron: _____
- F. A Chlorine Atom That Has Gained One Electron: _____

3. Write The Symbol For The Isotopes:

- A. With 8 Protons and 8 Neutrons
- B. With 16 Protons and 17 Neutrons
- C. With 28 Protons and 30 Neutrons
- D. With 92 Protons and 146 Neutrons

4. Here are the three isotopes of hydrogen. Write in below each symbol how many neutrons that isotope has.



5. Carbon has three isotopes. One has 6 neutrons, one has 7 neutrons, and one has 8. Draw their symbols & structures for these isotopes.